

BRODSKIY, A.I.; POMENKO, A.S.; ABRAMOVA, T.M.; FURMAN, Ye.G.; DARYEVA,  
E.P.; KUCHTENKO, I.I.; GALINA, A.A.

Electron paramagnetic resonance spectra of radicals arising in  
X-raying of polyamides. Dokl. AN SSSR 156 no. 5:1147-1149  
Je '64. (MIRA 17:6)

1. Institut fizicheskoy khimii im. L.V.Pisarzhevskogo AN  
UkrSSR. 2.Chlen-korrespondent AN SSSR (for Brodskiy).

311c

L 32715-65 EWD(j)/EWT(m)/EPF(c)/EPF(n)-2/T/EWP(j)/EWA(1)/EWA(h) Ec-L/Pr-L/  
Feb/Fu-L GG/RM

ACCESSION NR: AP5003833

S/0190/65/007/001/0116/0122

AUTHORS: Brodskiy, A. I.; Fomanko, A. S.; Abramova, T. M.; Danyeyev, E. B.;  
Galina, A. A.; Furman, Ye. G.; Kotorlenko, L. A.; Gardenina, R. A. 1/2  
40  
B

TITLE: Study of the radiation oxidation of poly-ε-caproamide 15

SOURCE: Vysokomolekulyarnyye soyedineniya, v. 7, no. 1, 1965, 116-122

TOPIC TAGS: caproamide, radiolysis, radiation oxidation, IR analysis 19

ABSTRACT: The nature of the radicals which occur in poly-ε-caproamide (PKA) during γ-radiation, the gaseous products of radiolysis and radiation oxidation, the changes in the IR spectra, changes in viscosity and content of amino end-groups, and the formation of hydroperoxides during radiation of PKA in oxygen were investigated. Films and slivers of PKA obtained by polymerization of ε-caprolactam in N<sub>2</sub> with H<sub>2</sub>O as initiator were used. The spectra of electron paramagnetic resonance (EPR) showed that the radical  $-\dot{\text{C}}\text{H}_2\text{CONHCH}_2-$  was formed (A. I. Brodskiy et al., Dokl. AN SSSR, 156, 1147, 1964). Chromatographic analysis of the gaseous products of radiolysis in vacuum and radiation oxidation in O<sub>2</sub> gave the H<sub>2</sub> and CO (with less than 3% CO<sub>2</sub>) separation shown in Fig. 1

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L 32715-65

ACCESSION NR: AP5003833

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on the Enclosure. The IR spectra of a PKA film (see Fig. 2 on the Enclosure) agree well with previous results of N. D. Slovokhotova (Dokl. AN SSSR, 127, 831, 1959). The effects of different radiation regimes on viscosity are shown in Fig. 3 on the Enclosure, the formation of  $\text{NH}_2$  and  $\text{COOH}$  end groups is given in Fig. 4 on the Enclosure, and the formation of peroxides is shown in Fig. 5 on the Enclosure. Orig. art. has: 6 figures.

ASSOCIATION: Institut fizicheskoy khimii im. L. V. Pisarzhevskogo AN UkrSSR (Physical Chemistry Institute, AN UkrSSR); Kievskiy filial Vsesoyuznogo nauchno-issledovatel'skogo instituta iskusstvennogo volokna (Kiev Section of the All-Union Scientific Research Institute of Synthetic Fibers)

SUBMITTED: 20Mar64

ENCL: 05

SUB CODE: CC

NO REF SOV: 004

OTHER: 000

Card 2/7

1C

L 24491-66 EPF(n)-2/EWT(m)/EWP(j)/T/EWA(h)/EWA(1) GG/RM/WW/JW  
 ACC NR: AP6006980 (A) SOURCE CODE: UR/0190/66/008/002/0261/0266

AUTHORS: Fomenko, A. S.; Abramova, T. M.; Dar'yeva, E. P.; Galina, A. A.; Furman, Ye. G.

ORG: Institute of Physical Chemistry in. L. V. Pissarshevskiy (Institut fizicheskoy khimii)

TITLE: Mechanism of action of di- $\beta$ -naphthyl-p-phenylenediamine during radiation oxidation of polycaproamide

SOURCE: Vysokomolekulyarnyye soyedineniya, v. 8, no. 2, 1966, 261-266

TOPIC TAGS: polyamide, free radical, oxidation kinetics

ABSTRACT: The effect of di- $\beta$ -naphthyl-p-phenylenediamine (I) upon the kinetics of accumulation of free radicals formed during the process of radiation-induced decomposition of polycaproamide (II) was investigated, and the yield of gaseous and oxygen-containing products of radiation-induced oxidation of the polymer was determined. The changes of the content of terminal  $\text{NH}_2$  groups, viscosity, and IR spectra occurring in II stabilized with I during the radiation-induced oxidation were also studied. The methods, involving ESR, chromatographic, chemical, and IR

Card 1/2

UDC: 678.01:54+678.675

L 24491-66

ACC NR: AP6006980

spectroscopic studies, were previously described by A. S. Fomenko, T. M. Abramova, E. P. Dar'yeva, A. A. Galina, and Ye. G. Furman (Ukr. khimich. zh., 30, 376, 1964). It was established that I has no effect upon breaking of the C-H, C-N, and -C-CO-bonds during radiolysis of II, but does affect C-N and C-CO bonds during radiation oxidation of II. The amount of peroxy carbonyl and carboxyl compounds formed during radiation oxidation of II stabilized with I is considerably lowered as compared with the untreated II. A possible mechanism for the inhibiting action of I is offered. Orig. art. has: 1 table, 6 figures, and 4 equations.

SUB CODE: 07, 11/ SUBM DATE: 05Mar65/ ORIG REF: 007

Card 2/2 88

L 40099-66 ENT(m)/ENP(j)/1 IJP(c) GG/RM

ACC NR: AP6019661

(A)

SOURCE CODE: UR/0073/66/032/006/0549/0554

AUTHOR: Brodskiy, A. I.; Fomenko, A. S.; Dar'yeva, E. P.; Abramova, T. M.; Galina, A. A.; Furman, Ye. G.

ORG: Institute of Physical Chemistry im. L. V. Pissarzhevskiy, AN UkrSSR (Institut fizicheskoy khimii AN UkrSSR)

TITLE: Gas evolution during the radiative-oxidative degradation of poly-ε-caproamide

SOURCE: Ukrainskiy khimicheskiy zhurnal, v. 32, no. 6, 1966, 549-554

TOPIC TAGS: polyamide, oxidative degradation, hydrogen, carbon monoxide, gamma radiation, radiation effect

ABSTRACT: Chromatographic analysis was used to find the radiation yields of hydrogen and carbon monoxide, the main gaseous products of the radiolysis and radiative oxidation of poly-ε-caproamide.  $G_{H_2}$  is about 1 mole/100 eV for both processes, and does not change as the dose rate increases from 0.4 to  $5.0 \times 10^{18}$  eV/g min.  $G_{CO}$  is equal to 0.3 mole/100 eV for radiolysis and to 0.6 mole/100 eV for radiative oxidation, and rises to 0.9 mole/100 eV as the dose rate increases from 0.4 to  $5.0 \times 10^{18}$  eV/g min. It was found that the combined action of gamma radiation and increased temperature approximately doubles the values of  $G_{H_2}$  and  $G_{CO}$  in both the radiolysis and radiative oxidation of poly-ε-caproamide in the case of a low dose rate of gamma radiation, and that the effect of this combined action on  $G_{H_2}$  and  $G_{CO}$  diminishes with increasing

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UDC: 678.01:54+678.675



L 40099-66

ACC NR: AP6019661

dose rate. It is shown that the stabilization of poly- $\epsilon$ -caproamide by the addition of the antioxidant di- $\beta$ -naphthyl-p-phenylenediamine does not change  $G_{H_2}$  during radiolysis and radiative oxidation, but markedly reduces the amount of carbon monoxide formed during radiative oxidation. Orig. art. has: 6 figures and 3 tables.

SUB CODE: 07/ SUM DATE: 31Jan64/ ORIG REF: 006

Card 2/2 *ML*

BASS-SHADKHAH, Kh.; GALINA. I.

Influence of some trace elements on the quantity and quality of zymosan formation in yeast. Report 1. Some chemical indices of zymosan obtained from yeast in a nutrient medium enriched with cobalt, copper, and manganese [with summary in English]. **Vestis Latv ak no.12:69-74 '61.**

1. AN Latviyskoy SSR, Institut eksperimental'noy i klinicheskoy meditsiny



*GALINA, I.G.*

USSR/General Problems of Pathology - Comparative Oncology.  
Tumors of Man.

U-3

Abs Jour : Ref Zhur - Biol., No 16, 1958, 75600

Author : Galina, I.G.  
Inst : -

Title : Rare Case of Hemorrhagic Hemangiomatosis.

Orig Pub : Materialy po bor'be so zlokachestvennymi opukholyami. Ufa,  
vyp. 8, 1956, 42-50.

Abstract : Hemorrhagic angiomatosis (HA) a rare hereditary diseases  
is observed with equal frequency in both sexes. The disease  
is characterized by congenital weakness of the mesenchy-  
mal system and capillaro-vascular net. Two stages of di-  
sease are distinguished; hemorrhagic and angiomatous.  
Treatment of HA (Osler's disease) remains unsolved. A ca-  
se of HA in a boy of 15 is described. The disease was er-  
roneously taken for rheumatic polyarthrititis and acute leu-  
kosis. After treatment with repeated blood transfusions

Card 1/2

MIKHEYEV, V.V.; SHTUL'MAN, D.R.; IL'YINA, N.A.; GALINA, I.V.; KOLOSOVA, O.A.

Amyotrophic lateral sclerosis syndrome in cervical osteochondrosis.  
Zhur. nevr. i. psikh. 63 no.6:833-840 '63. (MIRA 17:6)

1. Klinika nervnykh bolezney (direktor - prof. V.V. Mikheyev)  
I Moskovskogo ordena Lenina meditsinskogo instituta imeni I.M.  
Sechenova.

VIKTOROV, I.T.; GALINA, L.S.

Aminazine therapy of mental patients suffering from pulmonary tuberculosis. Vop.psikh.i nevr. no.7:293-304 '61. (MIRA 15:8)

1. Iz psikhiatricheskoy bol'nitsy imeni P.P.Kashchenko.  
(TUBERCULOSIS) (MENTAL ILLNESS) (CHLORPROMAZINE)

ZOZULYA, M.; (GALINA, V.

Metal workers who are interested in radio. Radio no.8:11-12  
Ag '60. (MIRA 13:9)  
(Radio clubs) (Radio operators)

GALINAT, A.

Experimental studies on the utilization of natural food by  
young crucian (*Carassius carassius* L.) and tench (*Tinca tinca* L.).  
Polskie arch hydrobiol 8:129-152 '61.

1. Zakład Ochrony Przyrody i Ekologii, Uniwersytet im. M.  
Kopernika, Toruń.

GALINICH, V.I., inzh.; KOLISNYK, V.N., inzh.; KOTANZHI, Yu.V., inzh.;  
OSOCHENKO, I.M., inzh.; SERGEYEV, I.I., inzh.

Using a slag crust for the production of AN-60 flux. Avtom.  
svar. 17 no.11:86-91 N '64 (MIRA 18:1)

1. Insti'tut elektrosvariki imeni Ye.O. Patona AN UkrSSR (for Galinich; Kolisnyk). 2. Khartsyzskiy trubnyy zavod (for Kotanahi, Osochenko). 3. Chelyabinskiy truboprokatnyy zavod (for Sergeyev).

2300

2808 1573 2515 2708

27029

S/125/61/000/002/001/013

A161/A133

AUTHORS: Galinich, V.I., Podgayetskiy, V. V.

TITLE: The effect of nitrogen on the porosity of welds in argon and carbon dioxide steel welding

PERIODICAL: Avtomaticheskaya svarka, no. 2, 1961, 24-32

TEXT: The purpose of the described experiments was to determine the shielding properties of different gas media against nitrogen. The welding tests were carried out in a sealed vacuum chamber with a rotating steel disk and sealed wire holder, using direct current and inverse polarity. The chamber was evacuated to  $10^{-2}$  mm Hg and filled with argon of different purity, oxygen and nitrogen, alimentary  $\text{CO}_2$ , and especially prepared carbon and nitrogen oxides. Commercial argon proved not suitable, for pores appeared in welds when more than 5%  $\text{N}_2$  was present in the gas. Oxygen added to such gas (argon with N) augmented the porosity. The data are different from those obtained by Ludwig (Ref. 3: H. C. Ludwig, Nitrogen effects in argon arc welding atmospheres, "The Welding Journal", no. 9, 1955, 4095-4145) who recommended argon with maximum 1%  $\text{N}_2$ , i.e. a higher N-content appears permissible. In combination with  $\text{CO}_2$ , already 1%  $\text{N}_2$  in gas caused

Card 1/3



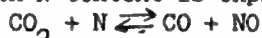
27029

S/125/61/000/002/001/013

A161/A133

The effect of nitrogen on the porosity ...

pores. The absorption of N by the weld metal from argon and from CO<sub>2</sub> was very different, and CO<sub>2</sub> is obviously no neutral solvent for N for it raises the solubility of N in the weld pool. Slag formed on the weld, and its composition was a proof of the oxidizing effect of CO<sub>2</sub>. Contrary to data of H. Schenck, G. Froberg and H. Graf [Ref. 5: "Archiv fuer das Eisenhuettenwesen", Heft 6, 329-337; (II)30 (1959)], oxygen had apparently no effect at all on the N absorption by the liquid metal. Metal deposited in a NO atmosphere was very porous. The more intensive N-absorption by the pool from CO<sub>2</sub> with N-content compared to argon with N-content is explained by the oxidizing reaction



where N - atomic nitrogen formed through dissociation of molecular nitrogen in the arc. The data show that the permissible N-content in argon is ten times that of the permissible N in CO<sub>2</sub>. The following conclusions are drawn: The higher the oxidizing effect of gas the more N is dissolved in the welding pool; the main cause of the rising nitrogen solubility in the welding pool with the increasing oxidizing properties of gas is the formation of gaseous NO, and the N content in the weld metal is higher after welding in NO than in N; the permissible N content in CO<sub>2</sub> is 0.1%; CO has an oxidizing as well as a

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27029

S/125/61/000/002/001/013

A161/A133

The effect of nitrogen on the porosity ...

carbonizing effect on the welding pool. There are 5 figures, 6 tables and 10 references: 5 Soviet-bloc and 5 non-Soviet-bloc. Three references to English, language publications read as follows: H. C. Ludwig, Nitrogen effects in argon arc welding atmospheres, "The Welding Journal", no. 9, 1955, 4095-4145; F. R. Hensel, Westinghouse Research Reports, R-74191, 1932; N. W. Krase, B. Mackey, Journal Phys. Chem. 1928 (32), 1488 (Gmelins Handbuch, B. 4, Stickstoff, 1936).

ASSOCIATION: Ordena Trudovogo Krasnogo Znameni Institut elektrosvariki im. Ye. O. Patona AN USSR (Electric Welding Institute "Order of the Red Banner of Labor" im. Ye. O. Paton AS UkrSSR)

SUBMITTED: July 4, 1960

Card 3/3

PODGAYETSKIY, V.V.; GALINICH, V.I.

Attainability of thermodynamic equilibrium in electric arc welding. Avtom. svar. 14 no.8:3-12 Ag '61. (MIRA 14:9)

1. Ordena Trudovogo Krasnogo Znameni Institut elektrosvarki imeni Ye.O. Patona AN USSR.

(Electric welding) (Thermodynamics)

GALINICH, V. I.

S/021/63/000/002/012/016  
D405/D301

GALINICH

AUTHORS: Pidhayeta'kyy, V. V. and Halynych, V. I.

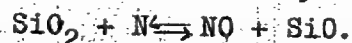
TITLE: Role of pre-electrode arc region in the absorption of nitrogen by the metal of a steel welding bath

PERIODICAL: Akademiya nauk UkrRSR. Dopovidi. no. 2, 1963, 209-210

TEXT: The experimental weldings were carried out in an argon atmosphere, containing 4% nitrogen. On welding by a direct current, the nitrogen is absorbed on the cathode only. In the case of an a.c. the nitrogen is absorbed by the liquid metal of the melted electrode, as well as by the metal of the welding bath, in accordance with the change in polarity. It was found that the absorption of nitrogen by the metal is enhanced by introducing solid oxide compounds into the welding region; the compound used was  $\text{SiO}_2$  powder. The corresponding chemical reaction can be written in the form:

Card 1/2

Role of pre-electrode ...

S/021/63/000/002/012/016  
D405/D301

The nitrogen oxides which are formed promote the absorption of nitrogen on the cathode as well as on the anode. Other experiments were conducted in an argon-nitrogen-hydrogen atmosphere, containing 0.015 - 0.016% N, and also in a  $\text{Ar} + \text{NH}_3$  atmosphere. It was found that the hydrogen promoted the absorption of nitrogen on the anode. Hence the introduction of oxygen and hydrogen in the arc atmosphere leads to the absorption of nitrogen. There are 3 tables.

ASSOCIATION: Instytut elektrozvaruvanniya AN URSSR (Institute of Electrical Welding of the AS UkrRSR)

PRESENTED: by Academician B. Ye. Paton

SUBMITTED: July 20, 1962

Card 2/2

PODGAYETSKIY, V.V.; GALINICH, V.I.

Effect of the kind and polarity of the current on nitrogen  
and hydrogen absorption by the welding bath. Avtom. svar.  
16 no.11:25-30 N '63. (MIRA 17:1)

1. Institut elektrosvariki imeni Ye.O. Patona AN UkrSSR.

LEVINTER, M.Kh.; IVANOVSKIY, G.F.; SMIRNOV, N.P.; GALIMOV, Zh.F.; GALINICH,  
Ye.T.

Remolding of catalytic cracking units using a spherical catalyst.  
Khim.i tekhn.topl.i masel 6 no.4:1-6 ap '61. (MIRA 14:3)

1. Upravleniye nerudnykh iskapayemykh i Novo-Ufimskiy nefteperera-  
baytvayushchiy zavod.  
(Cracking process) (Catalysts)



LEVINTER, M.Kh; IVANOVSKIY, G.F.; SMIRNOV, N.P.; GALINOV, Zh. F.;  
GALINICH, Ye.T.; GIMAYEV, R.N.

Modernization of catalytic cracking units at the Novoufinka  
Petroleum Refinery. Khim. i tekhn. topl. i masel 6 no. 7:1-6  
Jl '61. (MIRA' 14:6)

1. Novo-Ufimskiy neftepererabatyvayushchiy zavod i  
Upravleniye nerudnykh iskopayemykh.  
(Novoufinka—Cracking process—Equipment and supplies)

GALINICH, Ye.T.

Selecting an aluminosilicate catalyst for heavy feed stock cracking.  
Khim.i tekhn. topl.i masl 7 no.2:18-19 F '62. (MIRA 15:1)

1. Novo-Ufimskiy neftepererabatyvayushchiy zavod.  
(Cracking process) (Aluminosilicates)

GALINICH, Ye.I.

Causes of the deactivation of a catalyst. Neft'ep. i neftekhim.  
no.6:14-15 '64. (MIRA 17:9)

1. Bashkirskiy nauchno-issledovatel'skiy institut po pererabotke nef'ti.

MORCHALOV, B.F.; GABRILOV, Ye.T.; LEVINER, N.Ye.; GIL'FEL'D, B.N.

Ways of reducing the consumption of catalysts in the cracking of heavy crudes. Khim. i tekhn. topl. i masel 10 no.9:14-17 3 '65.

(NIFA 18:9)

1. Ufimskiy neftyanoy nauchno-issledovatel'skiy institut i Bashkirskiy nauchno-issledovatel'skiy institut po pererabotke nefti.

GALININ, K.P., inzh.

Modernizing a stonecutter for the underground production of ashlar.

Mekh.stroi. 14 no.9:15-17 S '57.

(MIRA 10:11)

(Stonecutters)

GALININ, E. D., LEONTOVICH, A. N., SVIRIDENKOV, E. A., SPORCHIKOV, V. N., CHIZHIKOVA, Z. A.

"Radiation properties of a ruby crystal laser."

The kinetics of generation at room temperature and low temperature (down to -165°C) and properties of radiation coherence in a ruby laser were investigated.

The report was presented to the 11th Conference on Luminescence (Molecular luminescence and luminescence analysis) Minsk, 10-15 Sept. 1962.

GALININ, N.F. Professor  
Military Medical Academy imeni S.M. Kirov, (Order of Lenin)

An article by B. NIKOL'SKIY stated that a conference was recently held in Yalta for the physicians of the military health resorts of the Crimea, at which were discussed the problems of the further development and introduction into practice of climatic healing and functional diagnosis. Three professors from the academy made speeches: I.T. KURSTIN\*, whose talk was entitled, "On the Physiociological Principles of Cortex-Visceral Therapy," N.F. Galinin\* spoke on "The Meaning of Radiation (from the sun) in the System of Health Resort Healing, Its organization and Methods of Calculation;" and D.A. VINIKUROV\* spoke on "The Therapeutic and Organizational Bases of Healing Physical Culture in Military Sanitoriums." (Krasnaya, Zvezda, Moscow, 9 Jul 1954)

SO: SUM No 239, 13 Oct. 1954



GALININ, H.F., prof.; SVIDERSKAYA, T.A., kand.med.nauk

Ultraviolet rays of the sun. Zdorov' 6 no.8:1-3 Ag '60,  
(MIRA 13:8)

1. Chlen-korrespondent AMN SSSR (for Galinin).  
(ULTRAVIOLET RAYS--PHYSIOLOGICAL EFFECT)

BAGDONAITE, A.; GALINIS, V.; JANKEVICIENE, R.; LEKAVICIUS, A.;  
NATKEVICIUTE-IVANAUSKIENE, M.; PIPINYS, J.; PURVINAS, E.;  
RIBOKAITE, R.; SNARSKIS, P.; STANCEVICIUS, A.; SARKINIENE, I.;  
ZIEMYTE, E., red.; ANAITIS, J., tekhn. red.

[Flora of the Lithuanian S.S.R.] Lietuvos TSR flora. Autoriu  
kolektyvas: A. Bagdonaite ir kiti. Vilnius, Valstybine politi-  
nes ir mokslines literaturos leidykla. Vol.2. 1963. 714 p.  
(MIRA 16:9)

1. Lietuvos TSR Mokslu Akademija, Vilna. Botanikos institutas.  
(Lithuania--Angiosperms)

EXCERPTA MEDICA Sec 8 Vol 12/1 Neurology Jan 59

347. MENTAL DISTURBANCES IN THE ACUTE STAGE OF POLIOMYELITIS  
(Russian text) - Galinis V. I. - ZH. NEVROPAT. I PSIKHIAT. 1958,  
18/7 (789-792)

Mental disturbances ('asthenia') were noted in 40 of the 58 patients studied. In some there were episodes of restlessness and anxiety or of delirium superimposed on the asthenic state. The episodes of delirium, without extreme excitability, usually preceded sleep. In the paralytic cases the mental disturbances were much more severe. Follow-up observations are necessary to ascertain whether such disturbances are reactions of functional type to the acute infection or are due to organic changes in the nervous system.

(L, 8)

*Detskoye psichiatricheskoye otdeleniye*

GALINIS, V.I.

Clinical aspects of neuropsychical disorders during the restorative and late periods following poliomyelitis. Vop.psikh.i nevr. no.7:66-74 '61. (MIRA 15:8)

1. Iz dispansernogo otdeleniya (zav. M.I.Shakin) psikhonevrologicheskogo instituta imeni V.M.Bekhtereva; (dir. - chlen-korrespondent Akademiya pedagogicheskikh nauk RSFSR prof. V.N.Myasishchev).  
(POLIOMYELITIS) (NEUROPSYCHIATRY)

GALINIS, V.I.

Appearance of schizophrenic syndromes in persons who have had  
poliomyelitis. Sbor. trud. Len. nauchn. ob-va nevr. i psikh.  
no.6:217-220 '79. (MIRA 13:12)

1. Iz d'spansernogo otdeleniya (sav. - starshiy nauchnyy sotrudnik  
M.I. Shakin) Psikhonevrologicheskogo instituta imeni V.M. Bekhtereva  
(direktor - chlen-korrespondent Akademii pedagogicheskikh nauk  
RSFSR prof. V.N. Myasishchev).  
(SCHIZOPHRENIA) (POLIOMYELITIS)

SHNAYDERMAN, S.Ya.; GALINKER, E.V.

Pyrocatechol complexes of uranyl. Zhur.neorg.khim. 7 no.2:279-  
283 F '62. (MIRA 15:3)  
(Uranyl compounds) (Pyrocatechol)

GALINKER, E.V.; FALEOLOG, Ya.N.

Surface recombination at the germanium - electrolyte interface.  
Elektrokhimiya 1 no.11:1311-1318 N '65. (MIRA 18:11)

1. Institut fizicheskoy khimii AN SSSR.



GALINKER, I.S.																									
COMMON ELEMENTS													PROCESSES AND PROPERTIES INDEX												
<p><i>eo</i></p> <p>The preparation of anhydrous aluminum chloride from Tachassow-jar clay. N. N. VORONIN AND I. S. GALINKER. J. Chem. Ind. (Moscow) 7, 143-9 (1930); Chem. Zentr. 1930, II, 783. Preliminary expts. with pure <math>Al_2O_3</math> formed to 2-mm. briquets with tar.</p> <p>showed that complete chlorination occurs at 650-800°. Clay contg. oxides of Fe, Ti, Ca, Mg, S, C and H was formed to similar briquets and the chlorination was studied at 800-850°. The chlorination becomes rapid at 800°, but only 45% of the Cl is used to chlorinate <math>Al_2O_3</math>; the rest serves to chlorinate <math>SiO_2</math>, <math>Fe_2O_3</math> and <math>TiO_2</math>. If the briquets are not tightly pressed some Cl is allowed to pass unreacted. Dilm. of the gas with air and the addn. of alkali chlorides to the clay do not improve the yields. <math>FeCl_3</math> cannot be removed completely by distn. The chlorination of clay will not be commercially economical as long as no large opportunities for the sale of <math>SiCl_4</math> are found. A. B.</p>																									
<p>450-550 METALLURGICAL LITERATURE CLASSIFICATION</p>																									

AL GALINKER,

a-1

**Cryoscopy and conductivity of halogen acids in ethyl alcohol-benzene solution.** I. S. GALINKER (Ukrain. Chem. J., 1937, 12, 212-220).—Association of EtOH in  $C_6H_6$  rises with increasing [EtOH], and is less in presence of HCl, HBr, or HI. The depression of f.p. of solutions of HX in  $C_6H_6$  falls with increasing [EtOH] to a min., corresponding with formation of 3EtOH.HX. In absence of EtOH the acids are monomeric in  $C_6H_6$ . Conductivity in the systems  $C_6H_6$ -EtOH-HX rises in the series  $X = Cl < Br < I$ . R. T.

GALINKER, I. S.

Galinker, I. S. - "The work of the Electrochemical Laboratory, (The group for the electrolysis of fused layers and galvanotechnology)", Trudy Vsesoyuz. in-ta solovoy promsti, Vol. V, 1949, p. 31-35.

SO: U-4631, 16 Sept. 53, (Letopis 'nykh Statey, No. 24, 1949).

GALINKER, I. S.

"Oxidation of Nitroxyl Chloride," Zhur. Prik, Khim. 22, No. 5, 1949. All-Union Inst. of Soda Ind. Kharkov.

GALINKER, I. S.

USSR/chemistry, Physical - Hydrates 21 Nov 51

"Equilibrium in Systems of Molten Salt - Water Vapor: Systems NaOH - H<sub>2</sub>O and KOH - H<sub>2</sub>O," I. S. Galinker, V. I. Korobkov, Kharkov Agr Inst 1951

"Dok Ak Nauk SSSR" Vol LXXI, No 3, pp 407-410

In a specially constructed stainless steel bomb, water vapor at variable pressures and temps has been dissolved in molten KOH and molten NaOH. Plots isotherms for molar quantities of H<sub>2</sub>O dissolved in the alkali at pressures ranging from 0 to 240 kg/sq cm. These isotherms permit the calcn of the bond energies between the mols of liquid water and

214714

molten alkali. Calcd the bond energy from the work of one mole of water being transferred to the melt in an isothermal reversible process. The cycle consists of 3 steps: I, work of the isothermal evapn of water; II, work of the isothermal expansion of steam from a pressure to vapor tension of the satd steam over water, to that equal to the vapor tension of the steam over the melt; III, work of condensation of the steam from the vapor phase to the melt. Then the max work is I minus II plus (particularly step II).

214714

GALINKER, I. S.

USSR/ Chemistry - Physical chemistry

Card 1/1

Pub. 22 - 22/49

Authors : Gavrish, M. L., and Galinker, I. S.

Title : Solubility of salts at high temperatures

Periodical : Dok. AN SSSR 102/1, 89-91, May 1, 1955

Abstract : Experiments were conducted to determine the solubility of AgCl, AgBr, AgJ, CuCl, CuBr and CuJ salts at temperatures ranging from 160 to 360°. The results regarding the solubility of the investigated halide salts of monovalent Ag and Cu are given. Six USSR references (1948-1953). Table; drawing.

Institution : The Kharkov Agric. Inst. im. V. V. Dokuchayev

Presented by : Academician I. I. Chernyaev, November 4, 1954

GALINKER, I. S.

✓ 2620

ELECTROCONDUCTIVITY OF CONCENTRATED  
AQUEOUS SOLUTIONS OF LiCl, NaCl, AND KCl AT  
HIGH TEMPERATURES, I. M. Rodnyanski and I. E.

Galinker. (Khar'kov V. V. Dokuchaev Agri. Inst.)

Doklady Akad. Nauk 108, 118-18 (1955) Nov. 1. (In Russian)

Investigations of electroconductivity of 1 to 3N salt  
solutions of LiCl, NaCl, and KCl up to 340°C are  
discussed, and tabulations are given. (R.V.J.)

GILIN, V.G.

18  
Validation of method of secondary aluminum  
recovery. N. I. Gilin, U. Ya. Beshanova, and V. G.  
Gilin. *Trudy Mepl.* 29, No. 7, 69-67 (1989). Various  
methods used in the recovery of Al from Al slag are dis-  
cussed. The retilling method is recommended. Further  
studies of filtration, centrifuging, and more suitable fluxes  
are recommended. I. B. G. (18)

E2

pro. RB  
adg



Pz linker, I. S.

Equilibria in alkali metal hydroxide-water systems. V. I. Kozlov and I. B. Galkin (V.I. Dokl. Akad. Nauk, Kharkov). *Chem. Abstr.* 10:1170-82 (1969); cf. *Chem. Abstr.* 36:462. The vapor pressure of alkali metal hydroxide solids, in the range of 250-420° was detd. in an autoclave provided with a window so that the vol. of the liquid phase could be observed and hence the concn. of the solid, be calcd. At 400-420° and atm. pressure the bases are practically anhyd. and become solns. with an increase in the mole fraction  $N_1$  of  $H_2O$ . The solubility for  $H_2O$  decreases in the order  $KOH \rightarrow NaOH \rightarrow LiOH$ . Plots of  $\log P/N_1$  vs.  $P/N_1$  (activity of  $H_2O$  vapor) are linear functions for  $LiOH-H_2O$  (obeying Henry's law) at 350, 375, and 420°. Henry's law is not followed by the systems  $KOH-H_2O$  and  $NaOH-H_2O$ . The calcd. heat capacities and the heats of diss. are tabulated. J. Heccowill

GALINKER, I. S. (Khar'kov)

"On the Properties of Aqueous Electrolyte Solutions at High Temperatures,"

Report presented at Conference on the Effect of Solvents on the Properties of  
Electrolytes, Khar'kov, 14-16 Oct '57.

Zhurnal Fizicheskoy Khimii, 1958, Vol 32, Nr 4, pp 960-962.

GALINGER, I.S.

"Power-Chemical Exploitation of Fuels with the Aid of a Solid Heat Carrier,"  
paper submitted for the 1st National Congress, Czechoslovak Scientific  
Technical Society for Fuel Utilization, Karlovy Vary, Czechoslovakia,  
12-17 May 58.

5(4)

AUTHORS:

Rodnyanskiy, I. M., Galinker, I. S.,  
Korobkov, V. I.

SOV/20-126-2-28/64

TITLE:

The Electric Conductivity of the Aqueous Solutions of Sodium Hydroxide at High Temperatures (Elektroprovodnost' vodnykh rastvorov yedkogo natra pri vysokikh temperaturakh)

PERIODICAL:

Doklady Akademii nauk SSSR, 1959, Vol 126, Nr 2, pp 327-329 (USSR)

ABSTRACT:

Short reference is first made to several earlier papers dealing with this subject, in which, according to the nature of the electrolyte and its concentration at various temperatures maxima of conductivity were found: For the salts of trivalent, bivalent, and univalent metals at 60°, 100-115°, and 280-300° respectively. It was of interest to investigate the further course of electric conductivity within the temperature range above 340°. However, the solution of this problem entails experimental difficulties as to the selection of the material for the electric insulation of the electrolytic cell and the hermetical sealing of the current supply lines. The chemical industry is in need of methods for the determination of electric conductivity at high temperatures

Card 1/3

The Electric Conductivity of the Aqueous Solutions  
of Sodium Hydroxide at High Temperatures

SOV/20-126-2-28/64

( $\sim 360^\circ$ ) and even for the most aggressive media, i.e. for basic lyes. The electrolytic cell used by the authors and the electrolytic conductors built into the steel stoppers of the autoclave are shown by a schematical drawing and briefly discussed. Next, the method of measuring electric conductivity is described. These measurements were carried out by means of the bridge MVL-47. A diagram shows the curves for the variation of the specific electric conductivity  $\kappa$  of aqueous NaOH solutions of various concentrations (1.3 and 5 %) up to  $360^\circ$ . All curves pass through a maximum near a temperature of  $200-220^\circ$  C. With increasing concentration the maximum shifts towards lower temperatures. At  $360^\circ$  the specific electric conductivity is by 2.5-3 times lower than maximum electric conductivity. The maximum of the conductivity for sodium hydroxide solutions is attained at lower temperatures than in the case of NaCl. At moderate temperatures NaCl and NaOH are equally strong electrolytes, but with increasing temperature, NaOH becomes a weaker electrolyte than NaCl. This is probably due to the existence of a larger

Card 2/3

The Electric Conductivity of the Aqueous Solutions of Sodium Hydroxide at High Temperatures SOV/20-126-2-28/64

portion of covalent binding in the molecule of sodium hydroxide. An exact interpretation of the process will be possible only after a large number of experimental data will have accumulated. There are 2 figures, 1 table, and 6 references, 5 of which are Soviet.

ASSOCIATION: Khar'kovskiy sel'skokhozyaystvennyy institut im. V. V. Dokuchayeva (Khar'kov Agricultural Institute imeni V. V. Dokuchayev)

PRESENTED: March 3, 1959, by A. N. Frumkin, Academician

SUBMITTED: February 9, 1959.

Card 3/3

GALINKER, I.S.; URAZOVSKIY, S.S.; BUDNIKOV, P.P.; KADANKER, L.I.;  
GORBANEV, A.I.

Andrei Nikitich Sysoev. Zhur.fiz.khim. 34 no.9:2130-2133 S  
'60. (MIRA 13:9)  
(Sysoev, Andrei Nikitich, 1901-1959)

S/076/60/034/011/021/024  
B004/B064

AUTHORS: Galinker, I. S. and Gavrish, M. L.

TITLE: Solubility of Alkaline-earth Hydroxides at High Temperatures

PERIODICAL: Zhurnal fizicheskoy khimii, 1960, Vol. 34, No. 11,  
pp. 2610-2612

TEXT: On the basis of their studies of the solubility of silver and copper halide compounds in water at temperatures of up to 340°C (Ref. 1), the authors arrived at the following conclusions: 1) Solubility increases rapidly with rising temperature in all readily melting salts which have a large number of covalent bonds. 2) In all salts with predominantly ionic binding and melting points above 800°C, the solubility first rises with temperature, and decreases above 300°C. These compounds become almost insoluble at the critical temperature of water. The authors have checked the validity of these assumptions by means of alkaline-earth hydroxides. The published data for  $\text{Ca(OH)}_2$  hold only up to 190°C; for  $\text{Sr(OH)}_2$ , up to 100°C; and for  $\text{Ba(OH)}_2$ , up to 109°C; solubility

Card 1/2



Solubility of Alkaline-earth Hydroxides at  
High Temperatures

S/076/60/034/011/021/024  
B004/B064

risers with temperature. According to the authors, especially the strongly ionically bound  $\text{Ba}(\text{OH})_2$  should show a decreasing solubility at higher temperatures. This was examined by means of an autoclave introduced into a nitrate-nitrite melt with a  $\text{Ba}(\text{OH})_2$  solution from which samples were filtered off. The solubility of barium hydroxide decreases rapidly above  $250^\circ\text{C}$  and is only 0.5% at  $360^\circ\text{C}$ . This observation is in accordance with the authors' assumption. There are 1 figure, 1 table, and 1 Soviet reference.

SUBMITTED: February 18, 1959

Card 2/2

STENDER, Vladimir Vil'gel'movich, prof., doktor tekhn. nauk. Prini-  
mali uchastiye: KSENZHEK, Oktavian Stanislavovich, dots.,  
kand. tekhn. nauk; RAZINA, Ninel' Fedorovna, dots., kand. tekhn.  
nauk; SAGOYAN, Leonid Nikolayevich, dots., kand. tekhn. nauk;  
SLUTSKIY, Iosif Zinov'yevich, dots., kand. tekhn. nauk; GALINKER,  
I.S., prof., otv. red.; TRET'YAKOVA, A.N., red.; TROFIMENKO, A.S.,  
tekhn. red.

[Applied electrochemistry] Prikladnaia elektrokimiia. Khar'kov,  
Izd-vo Khar'kovskogo gos.univ. im. A.M.Gor'kogo, 1961. 538 p.  
(MIRA 15:6)

(Electrochemistry)

RODNYANSKIY, I.M.; KOROBEKOV, V.I.; GALINKER, I.S.

Contraction of aqueous solutions of alcohols at 237°C. Izv.vys.  
ucheb.zav.; khim.i khim.tekh. 5 no.1:62-64 '62. (MIRA 15:4)

1. Khar'kovskiy sel'skokhozyaystvennyy institut imeni V.V.  
Dokuchayeva, kafedra fizicheskoy khimii.  
(Alcohols)

S/081/62/000/019/025/053  
B144/B180

AUTHORS: Gavrilsh, M. D., Galickier, I. S.

TITLE: Behavior of melted magnesium oxide in water at high temperatures

PERIODICAL: Referativnyi zhurnal.. Khimiya, no. 19, 1962, 353, abstract  
19K190 (Tr. Khar'kovsk. s.-kh. in-ta, v. 35 (72), 1961, 83-86)

TEXT: Tests showed that MgO sintered at 2800°C remains practically undissolved after prolonged boiling in distilled water, and the surface of the sample is unaltered. If this lead-burned MgO is heated between 175 and 350°C, lumps are quickly quenched to  $Mg(OH)_2$  by water or by saturated steam. ✓

The chief factor in this process is the additional energy of the steam, under pressure. [Abstracter's note: Complete translation.]

Card 1/1

MEDVEDEV, P. I.; GALINKER, I. S.

Investigation of titanium dioxide hydrosols by the cryoscopic  
method. Koll. zhur. 24 no.6:717-720 N-D '62.  
(MIRA 16:1)

1. Khar'kovskiy sel'skokhozyaystvennyy institut, kafedra nauch.  
cheskoy khimii.

(Titanium oxide) (Colloids) (Cryoscopy)

RODNYANSKIY, I.M.; KOROBEKOV, V.I.; GALINKER, I.S.

Specific volumes of electrolyte solutions at high temperatures.  
Zhur.fiz.khim. 36 no.10:2216-2219 0 '62. (MIRA 17:4)

1. Khar'kovskiy sel'skokhozyaystvennyy institut imeni Dokuchayeva.

LEKAKH, N.B.; GALINKER, I.S.

Nature of titanyl sulfate solutions. Lakokras. mat. i ikh prim.  
no.5:24-26 '63. (MIRA 16:11)

1. Khar'kovskiy sel'skokhozyaystvennyy institut im. V.V.  
Dokuchayeva.

GALINKER, I.S.; BELOVA, N.A.

Thermal properties of aqueous solutions of cobaltous chloride in  
the range of temperature up to 300°. Zhur.ob.khim. 33 no.10:3119-  
3121 0 '63. (MIRA 16:11)



GAVRISH, M.L.; GALINKER, I.S.

Reduction of cupric chloride by metallic silver in aqueous  
solutions at high temperatures. Zhur.fiz.khim. 37 no.2:  
463-464 F '63. (MIRA 16:5)

1. Sel'skokhozyaystvennyy institut imeni Dokuchayeva.  
(Copper chlorides) (Silver) (Reduction, Electrolytic)

GAVRISH, M.I.; GALINKER, I.S.

Complex salts of silver and copper halides in aqueous  
solutions at 300°C. Zhur. neorg. khim. 9 no.5:1289-

1290 My '64.

(MIRA 17:9)

GALINKER, I.S.; RODNYANSKIY, I.M.; KOROBKOV, V.I.; LEKAKH, N.B.

Temperature-dependent differences in the thermodynamic properties of water and electrolyte solutions. Ukr. fiz. zhur. 9 no.4:401-405 Ap '64. (MIRA 17:8)

1. Sel'skokhozyaystvennyy institut im. V.V. Dokuchayeva, Khar'kov.

KRASNOSEL'SKIY, V.N.; BOUYEVA, E.Y.; I.M.; SHEYN, S.M., GALINEER, I.G.

Conductometric analysis method for the control of alkali melting  
of the salts of aromatic sulfo acids. Khim. prom. 41 no.5:384-  
385 My '65. (MIRA 18:6)

1. Rubezhanskiy filial Nauchno-issledovatel'skogo instituta  
organicheskikh poluproduktov i krasiteley.

LEKAKH, N.B.; GALINKER, I.S.

Hydrolysis of aqueous solutions of ferrisulfate at high  
temperatures. Zhur. fiz. khim. 39 no.6:1318-1320 Ja '65.  
(MIRA 18:11)

1. Khar'kovskiy sel'skokhozyaystvennyy institut. Submitted  
March 15, 1963.

*Galvanic Elements and a Range of Potentials of Metals in Non-Aqueous Solutions.* V. S. Galilsker and I. L. Katsnel'son (*Zhur. Obshch. Khim.*, 1945, 18, 711-717; *C. Abs.*, 1946, 68, 7013).—[In Russian]. Measurements of e.m.f. were made for the system: standard electrode/junction electrolyte/aluminium bromide + benzene + metal halide/metal. The standard electrode consists of a small silver sheet in a benzene solution containing 0.25 mol. aluminium bromide and 0.025 mol. silver bromide per mol. of benzene. The junction electrolyte contained 0.25 mol. aluminium bromide per mol. of benzene and was saturated with calcium chloride. The e.m.f. with reference to the standard electrode was measured for 5 cells in which the metal halide and metal were: cuprous iodide + copper, thallium bromide + thallium, antimony bromide + antimony, bismuth bromide + bismuth, and potassium iodide + aluminium. The e.m.f. for these cells was from - 0.625 to + 0.497 V., and increased in the order thallium, aluminium, copper, bismuth, and antimony. This order was the same as that observed for the displacement of the metals from their halides in benzene solutions. The magnitude of the e.m.f. and order of the metals differed from those observed in aqueous systems. The e.m.f. was also measured for cells made from 2 half-cells containing standard electrodes, but having the silver bromide varied from 0.012 to 0.507 mol. per mol. of aluminium chloride. The half-cell with the lower silver bromide concentration was the anode. Potentials as high as 0.017 V. were observed. The results suggest that the silver complex in these solutions has considerable stability.

ASB-3LA METALLURGICAL LITERATURE CLASSIFICATION

GALINKER, V. S.

Galinker, V. S. "On the acidity of aluminosilicate gels,"  
Izvestiya Kiyevsk. politekhn. in-ta, Vol. VIII, 1949 (on cover;  
1949), p. 153-56, - Bibliog: 6 items

SO: U-5241, 17 December 1953, (Letopis 'Zhurnal 'nykh Statey, No. 26, 1949)

GALINKER, V. S.

Galvanic cells in toluene. V. S. Galinker, Zhur.  
Obshch. Khim. (J. Gen. Chem.) 1956, 30, 1446.  
E.m.f. was measured between an electrode of a metal  $M$   
in a soln. of the bromide of  $M$  in dry PhMe +  $AlBr_3$  at the  
ratios  $AlBr_3/PhMe = 0.25$  and  $M$  bromide  $AlBr_3 = 0.1$   
against a Ag electrode in  $AlBr_3/PhMe = 0.25$ ,  $AgBr/$   
 $AlBr_3 = 0.1$ . Stable potentials were attained only after  
hr.; they remained unchanged for a day with Cu and Bi,  
but were less stable with As, Sb, and Al. The potentials  
measured (against the above Ag electrode) were:  $Al/Al^{+++}$   
 $-0.075$  v.;  $Cu/Cu^+$   $-0.026$ ;  $Ag/Ag^+$   $0.000$  (by defini-  
tion);  $Bi/Bi^{+++}$   $+0.309$ ;  $Sb/Sb^{+++}$   $+0.479$ ;  $As/As^{+++}$   
 $+0.530$  v. The order of potentials in PhMe is Al, Cu,  
Ag, Bi, Sb, As, as in  $C_6H_6$  (C.A. 40, 7013), different from  
the order Al, Bi, Sb, As, Cu, in aq. soln. N. Tson

Lab. Physical Chem., Kiev Polytech. Inst.



Galinker, V. S.

Physicochemical study of  $CdI_2$  solutions in mixed solvents. ch  
 V. S. Galinker (Polytech. inst., Kiev), *Ukrain. Khim. Zhurn.* 28: 626-30 (1954) (in Russian).—Elec. cond. and viscosity of  $CdI_2$  solns. in binary mixts. of liquids, all of which contain  $Me_2O$ , the 2nd component being  $MeOH$ ,  $EtOH$ ,  $PrOH$ , or  $BuOH$ , were detd. Mol. elec. cond. of  $CdI_2$ - $Me_2O$ - $MeOH$  soln. at various dilns. passes through a max. that is ascribed to a formation of  $CdI_2$ - $MeOH$  compd. Temp. coeff. of elec. cond. was detd. between 20 and 30°: neg. for  $CdI_2$  in  $Me_2O$ , pos. for  $CdI_2$  in alcs., and corresponds to zero for correlation intervals in  $Me_2O$ -alc. system. The intensities of decompn. of  $CdI_2$  soln. in  $Me_2O$  and various alcs., and of  $MeOH$ , are not too different (cf. Jones and Mahin, *C.A.* 4, 403).  
 Elisabeth Barabash

AD-7751

Galin Ker, V. S.

Electrochemical investigation of certain binary systems in nitromethane. V. S. Galin Ker and A. E. Fil's (Polytechnic Inst., Kiev). *Zhur. Khim. Khim. Khim.* 20, 403-410 (1955). *Chem. U.S.S.R.* 25, 433-440 (1955) (Eng. transl.). The soly., elec. cond., concentration, potential, and electrolytic were investigated in the binary systems  $\text{NaCl-AlCl}_3$ ,  $\text{KCl-AlBr}_3$ ,  $\text{KBr-AlBr}_3$ , and  $\text{KBr-SbBr}_3$  in  $\text{CH}_3\text{NO}_2$  (1% soln.). The specific elec. cond. of the Al and Sb halides increased on adding alk. metal halides. The halides ceased dissolving when their mol. ratio to the Al halide reached 1. That, together with the elec. cond. in the formation of complexes  $\text{NaCl-AlCl}_3$ ,  $\text{NaCl-AlBr}_3$ ,  $\text{KBr-AlBr}_3$ , and  $\text{KBr-SbBr}_3$  in 1% soln. On electrolysis of these solns., the Al halide complexes liberated  $\text{Na}$  or  $\text{K}$  at the cathode while the Sb complex liberated  $\text{Sb}$ . B. M. (1955)

*GALINKER V.S.*

USSR/Physical Chemistry. Electrochemistry.

B-12

Abs Jour : Ref Zhur - Khimiya, No 7, 1957, 22486.

Author : V. S. Galinker.

Inst : Not given

Title : Electric conductivity and viscosity of aluminum chloride solution in nitromethane.

Orig Pub : Zh. obshch. khimii. 1956, 26, No 6, 1564-1568.

Orig Pub : Specific electric conductivity  $\kappa$ , viscosity  $\eta$ , and specific weight of  $AlCl_3$  solutions in nitromethane is measured with concentrations of  $AlCl_3$  from 5 to 45% at 20, 30, and 40°. Though each of the solution components separately is almost not current conductive, the resulting solutions are good conductors:  $\kappa$  reaches at C-23% its maximum value ( $\sim 1.5 \cdot 10^{-2} \text{ ohm}^{-1} \text{ cm}^{-1}$ ). Passage of the curve ( $\kappa, c$ ) through its maximum is connected with a sharp increase of  $\eta$  at  $c > 23\%$ , inasmuch as the product  $\kappa\eta$  grows smoothly with the increase of C. Temperature coefficients of  $\kappa$  are negative under the maximum point and positive after it. The calculated molecular electric conductivity (counting on  $AlCl_3$ ) grows with dilution of solutions as is the case with typical electrolytic solutions.

Card 1/1

-162-

GALINKER, V.S.

Use of niobium as an indicator electrode in potentiometric analysis.  
Izv. vys. ucheb. zav.; khim. i khim. tekhn. 4 no. 2:189-192 '61.  
(MIRA 14:5)

1. Kiyevskiy politekhnicheskoy institut. Kafedra fizicheskoy i  
kolloidnoy khimii.  
(Niobium) (Potentiometric analysis)

MESHKOVA, L.S.; GALINKER, V.S.; KUDRA, O.K.

Lead plating with trilon B. Mashinostroenie no.3:77-79 My-Je '62.  
(MIRA 15:7)

(Lead plating)

TYAGAY, V.A.; GALINKER, V.S.; FENERLI, G.N.

Systems  $\text{CdCl}_2$  -  $\text{MCl}$  -  $\text{H}_2\text{O}$  based on electric conductance data.  
Zhur.neorg.khim. 7 no.5:1154-1158 My '62. (MIRA 15:7)

1. Kiyevskiy politekhnicheskiy institut.  
(Cadmium compounds--Electric properties)  
(Systems (Chemistry))

GALINIER, V.S.

Titanium as an indicator electrode in neutralization reactions.  
Zhur.anal.khim. 17 no.5:556-559 Ag '62. (MIRA 16:3)

1. Polytechnical Institute, Kiev.  
(Electrodes, Titanium) (Neutralization)

GALINKER, V. S.; TYAGAY, V. A.; FENERLI, G. N.

Viscosity of mixtures of aqueous solutions of electrolytes.

Zhur. fiz, khim. 36 no.12:2638-2643 D '62.

(MIRA 16:1)

1. Kiyevskiy politekhnicheskij institut.

(Electrolyte solutions) (Viscosity)



ACCESSION NR: AP4018067

S/0080/64/037/002/0342/0348

AUTHORS: Galinker, V.S.; Sapry\*kin, A.I.

TITLE: Electrodeposition of cadmium-zinc alloy from an electrolyte containing Trilon B as complexing agent.

SOURCE: Zhurnal prikladnoy khimii, v.37, no.2, 1964, 342-348 .

TOPIC TAGS: cadmium zinc alloy, electrodeposition, Trilon B, cathodic deposition, corrosion resistance, throwing power, nontoxic electrolyte, stable electrolyte

ABSTRACT: The possibility of using Trilon B as the complexing agent in electrodepositing Cd-Zn alloys to obtain a more stable and less toxic electrolyte than with cyanides was investigated. Some factors affecting the electrodeposit are: by increasing the Zn content of the electrolyte or the NaOH concentration, or current density the Zn content in the cathodic deposit is increased; by increasing Trilon B concentration in the electrolyte the Cd content of the cathodic deposit is increased somewhat; and with increasing temperature the Cd content increases rapidly until 60°C when the deposit is pure Cd. The cathodic

Card 1/2

ACCESSION NR: AP4018067

current yield increases somewhat with increasing temperature. An investigation of the possibility of using a combination anode consisting of parallel adjacent Zn and Cd plates showed that the current is unevenly distributed. In order to form an electrolyte of constant composition to obtain a Cd-Zn alloy (20% Zn), the ratio of the surfaces of the anode Cd and Zn plates should be 2:1. The following electrolyte is recommended to obtain a cathodic deposit consisting of about 20% Zn and 80% Cd (an alloy having corrosion resistance equivalent to that of pure Cd):  $\text{CdSO}_4$ --35-40,  $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$ --12-14, Trilon B--100-110 and NaOH 75-85 gm./l; anode and cathode current density of 1-1.5 amps/dm<sup>2</sup>, 20C, no agitation. The electrolyte has good throwing power and gives a dense light gray deposit with good adhesion to iron. Orig. art. has: 6 figures and 2 tables.

ASSOCIATION: None

SUBMITTED: 10May62

DATE ACQ: 19Mar64

ENCL: 00

SUB CODE: CH, ML

NR REF SOV: 012

OTHER: 007

Card 2/2

GALINKER, V.S.; SAPRYKIN, A.I.

Electrodeposition of a cadmium-zinc alloy from an electrolyte  
containing trilon B as a complex-forming agent. Zhur. prikl.  
khim. 37 no.2:342-348 F '64. (MIRA 17:9)

GALINKER, V.S.; SAMOYKIN, A.I.

Study of the electroplating of a lead-antimony alloy using the  $Sb^{124}$   
radioactive indicator. Ukr. khim. zhur. 31 no.6:578-584 '65.

(MIRA 18:7)

1. Kiyevskiy politekhnicheskii institut.

GALINKER, V.S.; MESHKOVA, L.S.; KUDRA, O.K.

Cathodic polarization during the separation of lead from trilonate  
complex electrolytes. Ukr. khim. zhur. 31 no.8:866-871 '65.

(MIRA 18:9)

1. Kiyevskiy politekhnicheskii institut.

GALINKER, V.S.

Use of additive values of viscosity in determining the composition of  
complexes in solutions. Zhur. fiz. khim. 39 no.3:802-803 Mr '65.  
(MIRA 18:7)

GALINKER, V.S.; MILOVZOROV, V.P.; KUDRA, O.K.

Study of a new electrolyte for copper electroplating. Ukr.  
khim. zhur. 31 no.9:957-951 '65. (MIRA 18:11)

1. Kiyevskiy politekhnicheskii institut.

GALINKIN, A.A.; FREYBERG, F.A.

Epidemiology of erysipeloid. Vest.ven.i derm. no.5:49 S-O '53.  
(MLRA 6:12)

1. Iz kliniki kozhnykh i venericheskikh bolezney Voronezhskogo  
meditsinskogo instituta.

(Erysipelas)



GALINKIN, A.A.

Trichosporosis nodosa. Vest.ven.i derm. no.2:50-51 Mr-4, '54.  
(MLRA 7:4)

1. Iz kliniki kozhnykh i venericheskikh bolezney (zaveduyushchiy kafedroy - professor A.K.Yakubson) Voronezhskogo meditsinskogo instituta (direktor - professor V.P.Radushkevich) i gorodskogo vendispensera (glavnyy vrach M.V.Kalinina). (Scalp--Diseases) (Medical mycology)

GALINKIN, A.A.

Fungi flora in Voronezh during the pre- and postwar years. Vest.  
ven. i derm. no.4:52-53 J1-Ag '55. (MLBA 8:12)

1. Iz Voronseshskogo kozhno-venerologicheskogo dispensera.  
(VORONEZH--FUNGI, PATHOGENIC)

GALINKIN, A.A. (Voronezh)

Industrial dermatitis caused by Schizandra chinensis tincture.  
Gig. truda i prof.zab. 5 no.6:47 Je '61. (MIRA 15:3)

1. Gorodskoy kozhno-venerologicheskii dispanser.  
(CHEMICAL WORKERS--DISEASES AND HYGIENE)  
~~(SCHIZANDRA--PHYSIOLOGICAL EFFECT)~~  
(SKIN--DISEASES)

GALINKIN, A.A. (Voronezh)

Occupational dermatitis caused by aminazine. Gig truda i  
prof. zab. 7 no.1:51-53 Ja'63 (MIRA 16:12)

1. Voronezhskiy meditsinskiy institut.

GALINKIN, A.I.; GUSEVA, M.V. (Voronezh)

Improve drug distribution to the public. Apt. delo 10 no.3:65-67  
My-Je '61. (MIRA 14:7)

(VORONEZH--DRUGSTORES)

GALINKIN

137-58-2-3547

Translation from: Referativnyy zhurnal, Metallurgiya, 1958, Nr 2, p 182 (USSR)

AUTHOR: Galinkin, B. Ye.

TITLE: Increasing the Corrosion Resistance of Sheet Iron by Alloying it with Small Amounts of Copper (Povysheniye korrozionnoy stoykosti listovogo chuguna putem legirovaniya yego malymi prisadkami medi)

PERIODICAL: V sb.: Novoye v liteynom proiz-ve. Voronezh, 1957, pp 42-47

ABSTRACT: The results of corrosion tests under atmospheric conditions and in 5-percent solutions of HCl, HNO<sub>3</sub>, NaOH, and NaCl of cuprous sheet iron (CSI), alloyed with 0.25% Cu and 0.04% Ti are presented. The tests show that the corrosion strength of CSI is higher than the corrosion strength of bulk production sheet by 40% in atmospheric conditions, by 300 percent in a 5% HCl solution, and 60% in a 5% NaCl solution on intermittent immersion. It is noted that CSI is suitable for use as a reliable substitute for roofing Fe. The positive effect of Cu on the corrosion strength of iron is explained by the mechanical protection of the compacted film on the corrosion products and also by the setting-in of anodic passivity. M. K.

Card 1/1

1. Iron--Corrosion--Effects of copper 2. Iron--copper alloys  
--Corrosion--Test results

137-58-3-5938

Translation from: Referativnyy zhurnal, Metallurgiya, 1958, Nr 3, p 212 (USSR)

AUTHOR: Galinkin, B. Ye.

TITLE: Modification by Magnesium Improves Mechanical Properties of Cast Iron Sheets (Uluchsheniye mekhanicheskikh svoystv listovogo chuguna posredstvom modifitsirovaniya yego magniyem)

PERIODICAL: V sb.: Novoye v liteyn. proiz-ve. Voronezh, 1957, pp 48-55

ABSTRACT: A sheet of spheroidized cast iron (SCI) is produced in a standard technological process in conjunction with modification by metallic Mg (0.5 percent of the weight of liquid metal). Compared with a standard sheet an SCI sheet on a ferrite base exhibits improved mechanical properties: a 30 percent increase in strength, 8-9 percent increase in hardness, a 150 percent increase in plasticity, and improved corrosion resistance. The production of SCI sheets offers numerous possibilities for their extensive utilization as roofing and structural materials in the place of steel sheets.

A. S.

Card 1/1

137-58-4-7737

Translation from: Referativnyy zhurnal, Metallurgiya, 1958, Nr 4, p 198 (USSR)

AUTHORS: Tavadze, F. N. , Galinkin B. Ye.

TITLE: Effect of Heat Treatment on the Corrosion Resistance of Iron  
(Vliyaniye termicheskoy obrabotki na korrozionnuyu stoykost' chuguna)

PERIODICAL: Tr. Gruz. politekhn. in-t, 1957, Nr 3 (51), pp 120-127

ABSTRACT: The results of a study of the effect of heat treatment (HT) in the 200-700°C temperature interval on the resistance of iron to corrosion (CR) are presented. Corrosion tests were made on gray iron with lamellar graphite, high-strength iron with spheroidal graphite, sheet irons with flake graphite, without Cu and with 0.25 percent Cu. The tests were run in 5 percent HCl, HNO<sub>3</sub>, NaOH, and NaCl solutions (with intermittent and constant immersion) and under atmospheric conditions. It was found that under full immersion in NaOH and NaCl solutions, and under atmospheric conditions, the CR of iron is independent of HT. HT also fails to affect the CR of iron containing Cu in all mediums. HT significantly affects the CR of iron only in solutions of acids and in NaCl (on alternating immersion and withdrawal). The CR of iron under these conditions

Card 1/2



137-58-4-7737

Effect of Heat Treatment on the Corrosion Resistance of Iron

increases as the tempering temperature or the isothermic treatment temperature of iron is increased to 450°, regardless of the form and nature of the graphite inclusions. Further increase in temperature of HT in the 500-700° interval induces a reduction in the CR of iron. Irons of martensitic structure have the highest CR.

M. K.

1. Iron--Corrosion--Temperature factors

Card 2/2

*GALINKIN, B. Ye.*

AUTHOR: None Given

129-58-8-15/16

TITLE: Dissertations (Dissertatsii)

PERIODICAL: Metallovedeniye i Obrabotka Metallov, 1958, Nr 8,  
p 63 (USSR)

ABSTRACT: For the Degree of Candidate of Technical Sciences:  
Babich, V. K. "Study of the Processes of Tempering of  
Hardened Steel and of Steel Deformed in the Cold State  
(Izucheniye protsessov otpuska zakalennoy i  
kholodnodeformirovannoy stali). Dnepropetrovsk, 1957,  
Dnepropetr. metallurg. in-t im. I. V. Stalina  
(Dnepropetrovsk Metallurgical Institute imeni I.V.Stalin);  
A. A. Vorob'yev, A. A. "Investigation of the New  
Technology of Strengthening by Surface Work Hardening"  
(Issledovaniye novoy tekhnologii uprochneniya  
poverkhnostey naklepom), Leningrad 1957, Leningr.  
politekhn. in-t (Leningrad Polytechnical Institute);  
B. Ye. Galinkin "Corrosion Stability of Cast Iron as a  
Function of Certain Methods of its Treatment"  
(Korrozionnaya stoykost' chuguna v zavisimosti ot  
nekotorykh metodov yego obrabotki), Voronezh, 1956, *LLB*  
Gruz. politekhn. in-t (Georgia Polytechnical Institute);

Card 1/6

*in SARKIN*

Dissertations

129-58-8-15/16

Dobrovol'skiy, S. I. "Elucidation of the Possibility of Studying the Stresses During Plastic Deformation by Illuminating Transparent Models with Polarised Light" (Vyyasneniye vozmozhnosti izucheniya napryazheniy pri plasticheskoy deformatsii putem prosvechivaniya prozrachnykh modeley polyarizovannym svetom), Minsk, 1957, AN SSSR, In-t metallurgii im. A. A. Baykova (Ac. Sc. USSR, Institute of Metallurgy imeni A. A. Baykov);  
G. F. Lepin "Investigation of Certain Relations of Creep and Relaxation Phenomena in Metals" (Issledovaniye nekotorykh zakonomernostey svyazi yavleniy polzuchesti i relaksatsii napryazheniy v metallakh), Moscow, 1957, AN SSSR, In-t metallurgii im. A. A. Baykova (Ac. Sc. USSR, Institute of Metallurgy imeni A. A. Baykov);  
Loginov, P. I. "Investigation of the Influence of Short Duration Over-loads of the Resonance Type on the Fatigue Strength of Structural Steel" (Issledovaniye vliyaniya kratkovremennykh peregruzok rezonansnogo tipa na ustalostnuyu prochnost' konstruktsionnoy stali), Leningrad, 1957, Leningr. politekhn. in-t im. M. I. Kalinina  
Card 2/6 (Leningrad Polytechnical Institute imeni M. I. Kalinin) ;

Dissertations

129-58-8-15/16

M. S. Polyak "Seeking of New High Speed Facing Alloys of Increased Stability and Establishment of a Rational Facing Technology" (Izyskaniye novykh bystrorezhushchikh naplavochnykh splavov povyshennoy stoykosti i ustanovleniye ratsionnal'noy tekhnologii ikh naplavki), Tbilisi, 1957, Gruz. politekhn. in-t im. S. M. Kirova (Georgia Polytechnical Institute imeni S. M. Kirov); Yu. A. Preobrazhenskaya "Micro-structural Deformation and Influence of the Deformations on the Heat Resistance" (Mikro-strukturnyye deformatsii i vliyaniye deformatsii na zharoprochnost') Moscow, 1957, Mosk. in-t tsvet. metallov i zolota im. M. I. Kalinina (Moscow Institute of Non-Ferrous Metals and Gold imeni M. I. Kalinin);

For the Degree of Candidate of Physico-Mathematical Sciences: Bykovskiy, Yu. A. "Investigation of the Photomagnetic Effects in Germanium" (Issledovaniye fotomagnitnykh effektov v germanii), Moscow, 1957, Mosk. inzh. fiz. in-t (Moscow Engineering-Phys. Institute);

Ye. L. Gal'perin "Changes of the Crystal Structure of Steel During Cold Treatment and During Heat Treatment"

Card 3/6 "Izmeneniye kristallicheskoy struktury stali pri kholodnoy

Dissertations

129-58-8-15/16

i termicheskoy obrabotke, Leningrad, 1957, Leningr.  
ped. in-t im. A. I. Gertsena. Kafedra obshchey fiziki  
(Leningrad Pedagogical Institute imeni A. I. Gertsen,  
Chair of General Physics);  
A. V. Grin' "Investigation of the Phenomena of Non-  
Elasticity in  $\alpha$ -solid Solutions of Aluminium with Magnesium"  
(Issledovaniye yavleniya neuprugosti v al'fa-tverdykh  
rastvorakh alyuminiya s magniyem), Sverdlovsk, 1957,  
AN SSSR, Ural'skiy filial (Ac. Sc. USSR, Ural Branch);  
D. N. Karlikov "Near Order and Viscosity of Liquid  
Amalgams of Cadmium and Zinc" (Blizhniy poryadok i  
vyazkost' zhidkikh amal'gam kadmiya i tsinka), Kiyev, 1957,  
Kiyevskiy gos. universitet im. T. G. Shevchenko  
(Kiyev State University imeni T. G. Shevchenko);  
O. G. Karpinskiy "Residual Stresses After Grinding of  
Metals" (Ostatochnyye napryazheniya posle shlifovaniya  
metallov), Moscow, 1957, Mosk. inzh.-fiz.in-t  
(Moscow Engineering-Phys. Institute);

For the Degree of Candidate of Chemical Sciences:

Card 4/6 I. K. Marshakov "Investigation of the Mechanism of Slot  
Corrosion of Metals" (Issledovaniye mekhanizma "shchelevoy

Dissertations

129-58-8-15/16

korrozii" metallov), Voronezh, 1957, AN SSSR. In-t fizicheskoy khimii (Ac. Sc. USSR, Institute of Physical Chemistry);

Molodova, K. A. "Complex Compounds of Platinum with Acetylene Derivatives in the Internal Sphere and Some of Their Properties" (Kompleksnyye soyedineniya platiny s atsetilenovymi proizvodnymi vo vnutrenney sfere i ikh nekotoryye svoystva), Leningrad, 1957, Leningr. ped. in-t im. A. I. Gertsena (Leningrad Pedagogical Institute imeni A. I. Gertsen);

B. I. Nabivanets "Study of the Complex Compounds of Mo in the Solution" (Izucheniye kompleksnykh soyedineniy molibdena v rastvore), Kiyev, 1957, AN Ukr.SSR, In-t obshchey i neorganicheskoy khimii (Ac. Sc. Ukr. SSR, Institute of General and Inorganic Chemistry);

G. A. Tedoradze "Study of the Kinetics of Oxidation of Chlorine Ions and Ionisation of Molecular Cl on platinum" (Izucheniye kinetiki okisleniya khlor-ionov i ionizatsii molekulyarnogo khloro na platine), Moscow, 1957, MGU im. M. V. Lomonosova. Kafedra elektro-khimii (Moscow State University imeni M.V.Lomonosov. Chair of Electro-chemistry);

Card 5/6

Dissertations

129-58-8-15/16

L. V. Petrova "Synthesis of  $\alpha$ -,  $\beta$ -unsaturated ketones,  $\beta$ -chlorketones, diketones and ketoxides in Presence of Metal Halogenides" (Sintez  $\alpha$ -,  $\beta$ -nepredel'-nykh ketonov,  $\beta$ -khlorketonov, diketonov i ketokislov, v prisutstvii galogenidov metallov), Moscow, 1957, AN SSSR, In-t organich. khimii im. N. D. Zelinskogo (Ac. Sc., USSR, Institute of Organic Chemistry imeni N. D. Zelinskiy)

1. Metallurgy--USSR

Card 6/6

S/081/61/000/024/039/086  
B117/B147

AUTHOR: Galinkin, B. Ye.

TITLE: Corrosion resistance of cold-worked cast-iron sheets

PERIODICAL: Referativnyy zhurnal. Khimiya, no. 24, 1961, 307, abstract 24I208 (Sb. "Polucheniye izdeliy iz zhidk. met. s uskoren. kristallizatsiyey", Moskva-Kiyev, Mashgiz, 1961, 297-298)

TEXT: The effect of cold rolling upon the corrosion resistance of cast-iron sheets is described. It is shown that the corrosion resistance increases with an increase of the degree of deformation up to 25%. This is related to the compacting and smoothing of the sheet surface. Further increase of the degree of deformation (up to 100%) of the sheet hardly improves its resistance corrosion. On the contrary, the corrosion resistance of such cast iron is even lower. To increase the corrosion resistance of cast-iron sheet, it is recommended that the rolling process should be performed at the optimum degree of deformation (25%) under the conditions of further recrystallization. [Abstracter's note: Complete translation.] ✓

Card 1/1